

Date Planned : __ / __ / __	Daily Tutorial Sheet-3	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	JEE Advanced (Archive)	Exact Duration : _____

31. The compound which gives off oxygen on moderate heating is : (1986)
(A) cupric oxide (B) mercuric oxide (C) zinc oxide (D) aluminium oxide
32. Arrange the following in increasing order of bond strength : HCl, HBr, HF, HI (1986)
33. Mention the products formed when : Chlorine gas is bubbled through a solution of ferrous bromide. (1986)
34. Mention the products formed when : Iodine is added to solution of stannous chloride. (1986)
35. Mention the products formed when : Sulphur dioxide gas, water vapour and air are passed over heated sodium chloride. (1986)
36. The bonds present in N_2O_5 are : (1986)
(A) only ionic (B) covalent and coordinate
(C) only covalent (D) covalent and ionic
- *37. In the electrolysis of alumina, cryolite is added to : (1986)
(A) lower the melting point of alumina (B) increase the electrical conductivity
(C) minimise the anode effect (D) remove impurities from alumina
38. Bromine can be liberated from potassium bromide solution by action of : (1987)
(A) iodine solution (B) chlorine water
(C) sodium chloride (D) potassium iodide
39. Amongst the trihalides of nitrogen which one is least basic ? (1987)
(A) NF_3 (B) NCl_3 (C) NBr_3 (D) NI_3
40. Which of the following oxides of nitrogen is a coloured gas ? (1987)
(A) N_2O (B) NO (C) N_2O_5 (D) NO_2
41. Write balanced equation for the reaction : Dilute nitric acid is slowly reacted with metallic tin. (1987)
42. Write balanced equation for the reaction : Phosphorous reacts with nitric acid to give equimolar ratio to nitric oxide and nitrogen dioxide. (1988)
43. Arrange the following in increasing order of thermal stability : $HOCl$, $HOClO_2$, $HOClO_3$, $HOClO$ (1988)
44. Arrange the following in increasing order of acidic character : CO_2 , N_2O_5 , SiO_2 , SO_3 (1988)
45. Give reason : Valency of oxygen is generally two whereas sulphur shows valency of two, four and six. (1988)